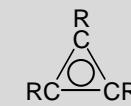
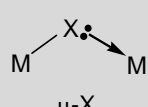
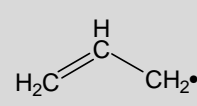

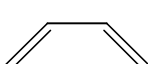

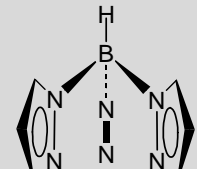


Table 1 Electron counting for commonly encountered ligands

<i>nVE</i>	<i>Ligands</i>
1-electron ligands	<p><i>Ligands which in the free state would be a radical:</i></p> <p>H, F, Cl, Br, I, OH, NR₂, OR, SR, CN, N₃, NCS, bent NO</p> <p>CR₃ – alkyl, aryl, alkenyl, alkynyl, formyl, acyl</p>
2-electron ligands	<p><i>Ligands which in the free state would have an even number of valence electrons:</i></p> <p>OH₂, NH₃, ethers, amines, thioethers, phosphines</p> <p>CO, CNR, CS, CR₂, C=CR₂</p> <p>Alkenes, alkynes; molecules which bind side-on through a multiple bond: O₂, SO₂, CS₂, RP=PR, R₂Si=CR₂</p>
3-electron ligands	<p>Linear NO, nitride (N)</p> <p><i>Ligands which can be subdivided into a combination of 1 VE and 2 VE donations:</i></p> <div style="display: flex; justify-content: space-around; align-items: center;"> <div style="text-align: center;">  <p>η³-cyclopropenyl</p> </div> <div style="text-align: center;">  <p>μ-X</p> </div> <div style="text-align: center;">  <p>η³-allyl</p> </div> </div>
4-electron ligands	<div style="display: flex; justify-content: space-around; align-items: center;"> <div style="text-align: center;">  <p>η⁴-cyclobutadienyl</p> </div> <div style="text-align: center;">  <p>diene</p> </div> </div>
5-electron ligands	<div style="display: flex; justify-content: space-around; align-items: center;"> <div style="text-align: center;">  <p>η⁵-cyclopentadienyl</p> </div> <div style="text-align: center;">  <p>tris(pyrazolyl)borate</p> </div> </div>
6-electron ligands	
7-electron ligands	
8-electron ligands	